

Supporting information: Sorting lithium-ion battery electrode materials using dielectrophoresis at frequencies of up to 500 kHz

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1 S1 - Signal processing

2 The data provided by the spectrometer was acquired by a Labview program every 250 ms
3 and stored in a text file. The files were then processed using a MATLAB script, which is
4 published in an online repository along with the measurement data.¹

5 First, the background intensity i_b is determined and subtracted from the signal (see
6 below). Second, the data was summed over several wavelengths (here, from 400 nm to
7 650 nm) and afterward smoothed using a moving average function to reduce the noise of the

8 data. The signal reduction sr was defined as the mean value of reduction of the intensity i
9 from 200 s (t_1) to 300 s (t_2) after the background was subtracted:²

$$sr = 1 - \frac{1}{t_2 - t_1} \int_{t_1}^{t_2} \frac{i - i_b}{i_{c_0} - i_b} dt. \quad (1)$$

10 All intensities in the study are normalized to the initial concentration i_{c_0} measured at the
11 beginning of the experiment.

12 **S2 - Background measurement**

13 The background signals were measured daily by using the medium without particles in it
14 for one minute. Before the background measurements were performed the signal needs to
15 be constant for several minutes. However, we found that this level changes over time to a
16 certain degree. For particles with high reflection or fluorescence, these changes are very small
17 compared to the signal and thus the effect is negligible. For the particles with lower reflection,
18 changes in the background can affect the signal reduction in the range of several percent.
19 The reason behind these fluctuations of the background signal can be of many origins. First,
20 the light source might vary over time in intensity. Second, the liquid light guide, which
21 connects light source and flow cell, could also show altered transmission behavior over time
22 due to increased temperature in it. Third, we observed that particles tend to sediment in
23 the flow cell, which is why we flushed the setup after two experiments. This might also
24 affect the background signal. Finally, the suspension contains dissolved gases from which
25 small bubbles can originate. These bubbles were observed to adhere to the flow cell. These
26 bubbles scatter the light and thus alter the signal recorded during the measurement. This
27 effect is minimized by degassing the solution prior to the experiment at ≈ 70 mbar.

28 S3 - Saturation effects

29 To investigate whether saturation effects of the filter at the used concentration occur, we
30 conducted experiments using Actilion at 50 V_{pp} and 500 kHz at 6 mLmin⁻¹. At this combi-
31 nation of parameters a complete removal of Actilion was observed. Now, the trapping time
32 was prolonged and monitored for almost four times longer compared to the experiments in
33 the manuscript (970 s vs. 270 s). The experiments did not show any saturation effects in
34 this time range (figure 1). The intensity was constant over the entire time range for all three
35 conducted experiments. After 1200 s the data acquisition was turned off and the channel
36 flushed at a high flow rate to prepare the channel for the next experiment with prolonged
trapping times.

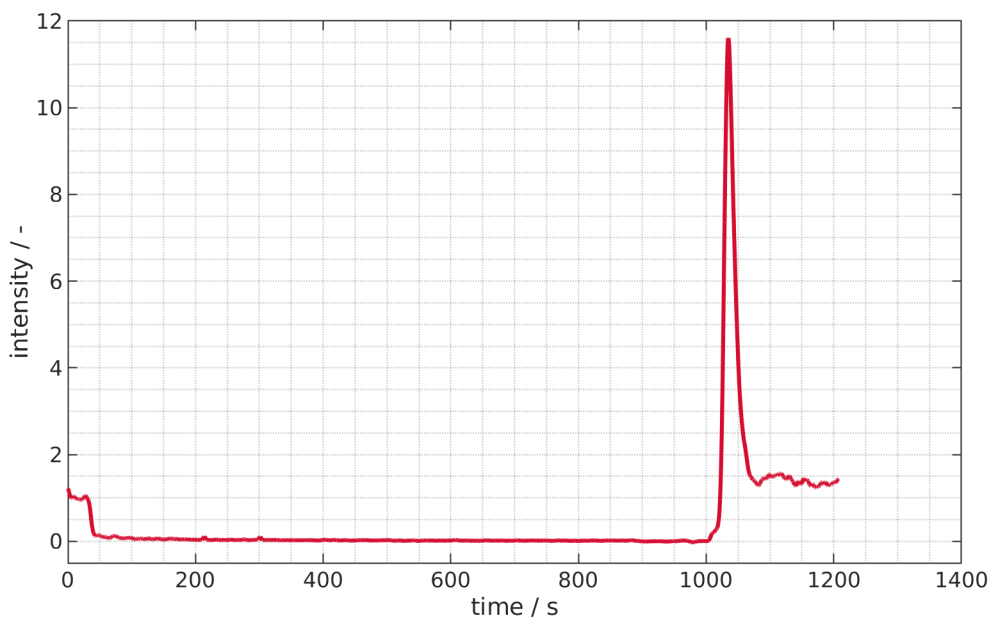


Figure 1: Intensity over time of Actilion at 50 V_{pp} and 500 kHz at 6 mLmin⁻¹. Voltage was turned on between 30 s and 1000 s.

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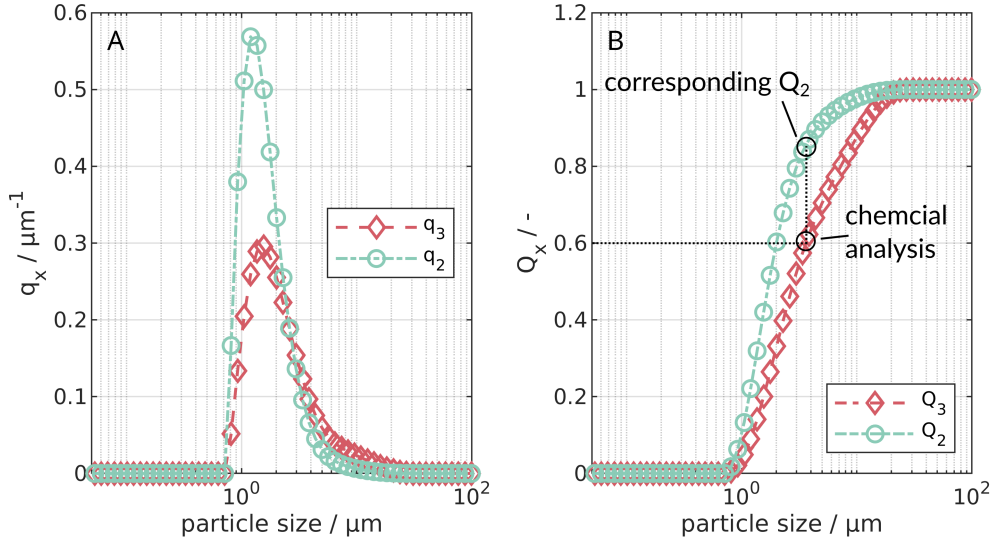


Figure 2: A: particle size distributions by surface (q_2) and volume (q_3) of LiFePO_4 microparticles. B: Cumulative particle size distributions by surface (Q_2) and volume (Q_3) of LiFePO_4 microparticles. Measured with a Mastersizer 2000. Measurement and calculations assume perfect spheres.

38 S4 - Effect of different particle distributions

39 To show the influence of different particle distributions, we measured the volume weighted
 40 particle size distribution q_3 of the LiFePO_4 particles (Mastersizer 2000, Malvern Panalytical
 41 GmbH, Germany) from which the cumulative distribution Q_3 can be calculated. The software
 42 of the measurement device assumes that the particles are spherical, which is only a rough
 43 approximation (figure 2 A&B, main document). Nonetheless, for illustration purposes this
 44 helps to demonstrate the effects of different particle size distributions. Please note, that there
 45 are deviations compared to table 1 in the main document. These might be due to different
 46 measurement techniques of the manufacturer and those presented here. Out of the measured
 47 q_3 distribution, the surface weighted particle size distribution q_2 can be calculated?

$$q_2(x) = \frac{x^{-1}q_3(x)}{\int_{x_{max}}^{x_{min}} x^{-1}q_3(x)dx}, \quad (2)$$

48 with x as the particle size and under the assumption of spherical particles. Figure 2
 49 A&B shows the different distributions. To explain the difference between reflection and

50 chemical analysis, figure 2 B is helpful. Assuming that first the larger particles are trapped
51 inside the channel and around 40 % trapping according to the chemical analysis, this cor-
52 responds to $Q_3 = 0.6$ (lower circle in figure 2 B). Transferring this particle size to the Q_2
53 distribution gives $Q_2 \approx 0.85$. This would lead to a reflection reduction of ≈ 15 %, which
54 is not too far off from the measured value of 19 ± 1 %. Obviously, this calculation in-
55 cludes assumptions but illustrates the intrinsic differences of chemical analysis and reflection
56 measurements quite clearly.

57 **S5 - SOP: Chemical analysis**

58 The determination of the total iron content with prior digestion, reduction to iron(II)-ions
59 and photometric determination was derived from DIN 38406 Part 1: Determination of iron
60 (German standard procedure for water, waste water and sludge analysis).³

61 The following points deviate from the DIN standard:

- 62 • All glass vessels were soaked in a Mucosal bath (about 0.2 % Mucosal) overnight.
63 Immediately afterward, the vessels were cleaned in a laboratory dishwasher using a
64 program that runs up to a maximum temperature of 55°C. Then they were soaked
65 overnight in 2 % nitric acid. The glass vessels were rinsed with pure water and allowed
66 to dry at room temperature.
- 67 • 4 mL sample was collected in a vessel and acidified. The sample was put into an
68 ultrasonic bath for 5 min, directly followed by pipetting 3.5 mL of the sample into
69 10 mL volumetric flask.
- 70 • 10 mL volumetric flasks were used for sample preparation and all volumes were adjusted
71 in the same ratio.
- 72 • Digestion was performed at room temperature directly in the 10 mL volumetric flask
73 by adding HNO₃ and HCl, mixing and allowing to stand for 2.5 h.

- 74 • Calibration solutions were prepared directly in 10 mL volumetric flask, then treated in
75 the same way as the samples.

Table 1: Calibration solution preparation in 10 mL volumetric flask

Iron concentration / mgL^{-1}	Volume of 10 mgL^{-1} Fe-standard / mL	Volume water/ mL
0	-	3.5
0.1	0.035	3.465
0.3	0.105	3.395
0.5	0.175	3.325
1	0.35	3.150
3	1.05	2.45
5	1.75	1.75

76 References

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78 trode Materials Using Dielectrophoresis at Frequencies of up to 500 kHz".
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